

UNIT – 3 – p BLOCK ELEMENTS -II**I. Answer in brief(2/3 marks)**

1. What is inter pair effect?
2. Chalcogens belongs to p-block. Give reason.
3. Explain why fluorine always exhibit oxidation state of -1?
4. Give the oxidation state of halogens in the following
 - a) OF_2
 - b) O_2F_2
 - c) Cl_2O_3
 - d) I_2O_4
5. What are interhalogen compounds? Give examples.
6. Why fluorine is more reactive than other halogens?
7. Give the uses of helium.
8. What is the hybridization of iodine in IF_7 ? Give its structure.
9. Give the balanced equation for the reaction between chlorine with cold NaOH and hot NaOH.
10. How will you prepare chlorine in the laboratory?
11. Give the uses of sulphuric acid.
12. Give a reason to support that sulphuric acid is a dehydrating agent.
13. Write the reason for the anomalous behavior of nitrogen.
14. Write the molecular and structural formula for the following molecules
 - a) Nitric acid
 - b) Dinitrogen Pentoxide
 - c) Phosphoric acid
 - d) Phosphine
15. Give the uses of argon.
16. Write the valence shell electronic configuration of group – 15 elements.
17. Give two equations to illustrate the chemical behavior of phosphine.
18. Give a reaction between nitric acid and a basic oxide.
19. What happens when PCl_5 is heated?
20. How will you prepare ammonia in the laboratory?

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II. Answer in a paragraph(5 marks)

1. Explain briefly the action of nitric acid on metals.
2. How are the following oxides of nitrogen prepared?
 - a) Nitrous oxide
 - b) Nitric oxide
 - c) Dinitrogen trioxide
 - d) Nitrogen dioxide
 - e) Nitrogen tetraoxide
 - f) Nitrogen pentoxide
3. How are the following oxyacid are of nitrogen prepared?
4. Explain the all-Allotropic form of phosphorus.
5. Explain the preparation, properties and uses of phosphorus trichloride.
6. Explain the preparation, properties and uses of phosphorus pentachloride
7. Explain the structure of phosphorous trichloride and phosphorus pentachloride.
8. Explain the Allotropic form of Sulphur.
9. Write the uses of Noble gases.